

THE PERIODIC TABLE

Period 1 2 3 4 5 6 7 0

1																	4		
																	He Helium 2		
2	7	9																	20
	Li Lithium 3	Be Beryllium 4																	Ne Neon 10
3	23	24																	35.5
	Na Sodium 11	Mg Magnesium 12																	Cl Chlorine 17
4	39	40																	80
	K Potassium 19	Ca Calcium 20																	Br Bromine 35
5	86	88																	127
	Rb Rubidium 37	Sr Strontium 38																	I Iodine 53
6	133	137																	210
	Cs Caesium 55	Ba Barium 56																	At Astatine 85
7	223	226																	222
	Fr Francium 87	Ra Radium 88																	Rn Radon 86

11	12	14	16	19	20
B Boron 5	C Carbon 6	N Nitrogen 7	O Oxygen 8	F Fluorine 9	Ne Neon 10
27	28	31	32	35.5	40
Al Aluminium 13	Si Silicon 14	P Phosphorus 15	S Sulfur 16	Cl Chlorine 17	Ar Argon 18
70	73	75	79	80	84
Ga Gallium 31	Ge Germanium 32	As Arsenic 33	Se Selenium 34	Br Bromine 35	Kr Krypton 36
115	119	122	128	127	131
In Indium 49	Sn Tin 50	Sb Antimony 51	Te Tellurium 52	I Iodine 53	Xe Xenon 54
204	207	209	210	210	222
Tl Thallium 81	Pb Lead 82	Bi Bismuth 83	Po Polonium 84	At Astatine 85	Rn Radon 86

65	63.5	59	59	56	55	52	51	48	45	44	43	42	41	39	38	37
Zn Zinc 30	Cu Copper 29	Ni Nickel 28	Co Cobalt 27	Fe Iron 26	Mn Manganese 25	Cr Chromium 24	V Vanadium 23	Ti Titanium 22	Sc Scandium 21	Ru Ruthenium 44	Rh Rhodium 45	Pd Palladium 46	Ag Silver 47	Cd Cadmium 48	Sr Strontium 38	Rb Rubidium 37
112	106	106	103	101	99	96	93	91	89	88	86	84	81	79	78	77
Cd Cadmium 48	Ag Silver 47	Pd Palladium 46	Rh Rhodium 45	Ru Ruthenium 44	Tc Technetium 43	Mo Molybdenum 42	Nb Niobium 41	Zr Zirconium 40	Y Yttrium 39	Ca Calcium 20	Sc Scandium 21	Ta Tantalum 73	Hf Hafnium 72	La Lanthanum 57	Ac Actinium 89	Ir Iridium 77
201	197	195	192	190	186	184	181	179	139	137	133	181	179	139	137	133
Hg Mercury 80	Au Gold 79	Pt Platinum 78	Ir Iridium 77	Os Osmium 76	Re Rhenium 75	W Tungsten 74	Ta Tantalum 73	Hf Hafnium 72	La Lanthanum 57	Ba Barium 56	Cs Caesium 55	Ta Tantalum 73	Hf Hafnium 72	La Lanthanum 57	Ba Barium 56	Cs Caesium 55

Key

Relative atomic mass
Symbol
Name
Atomic number



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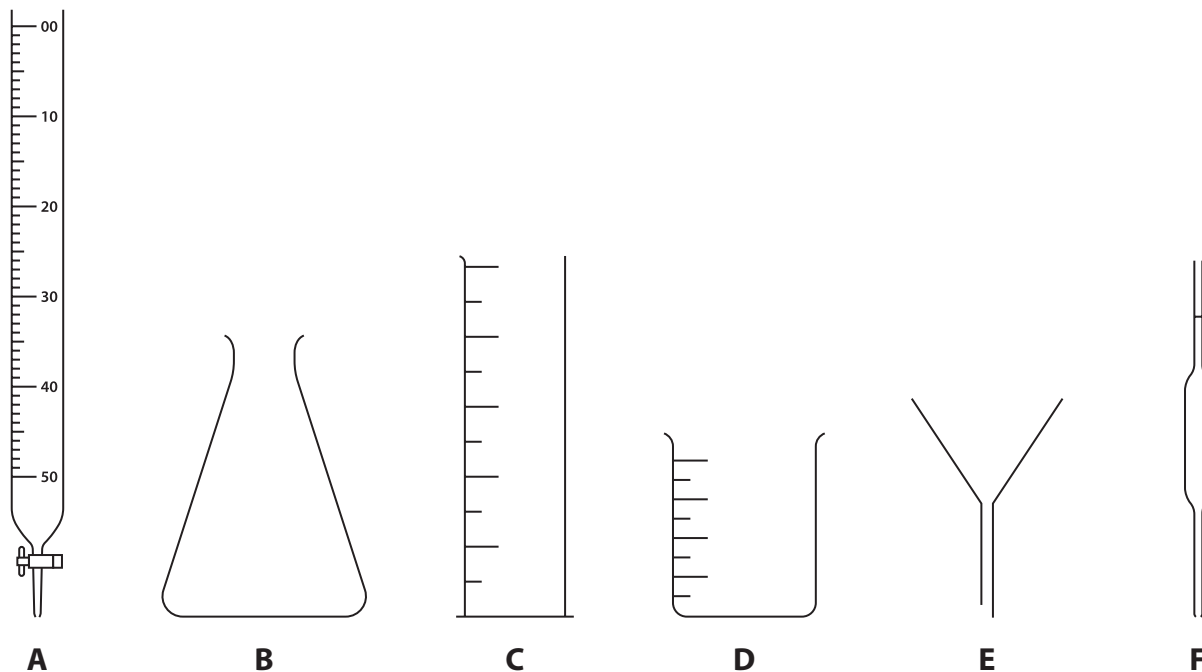
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Answer ALL questions.

1 The diagram shows six pieces of apparatus that are used in the laboratory.



The table lists the names of four pieces of apparatus.

Complete the table by giving a letter, A, B, C, D, E or F, to identify each piece of apparatus listed.

(4)

Name of apparatus	Letter
beaker	
burette	
measuring cylinder	
pipette	

(Total for Question 1 = 4 marks)

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2 Rubidium is an element in Group 1 of the Periodic Table.

A sample of rubidium contains two isotopes, ${}_{37}^{85}\text{Rb}$ and ${}_{37}^{87}\text{Rb}$

(a) (i) State how the nuclei of the two isotopes are similar. (1)

.....

.....

(ii) State how the nuclei of the two isotopes are different. (1)

.....

.....

(iii) How many electrons are in the outer shell of a rubidium atom? (1)

- A 1
- B 3
- C 9
- D 37

(b) The relative abundances of the two isotopes in the sample of rubidium are



Calculate the relative atomic mass of rubidium.

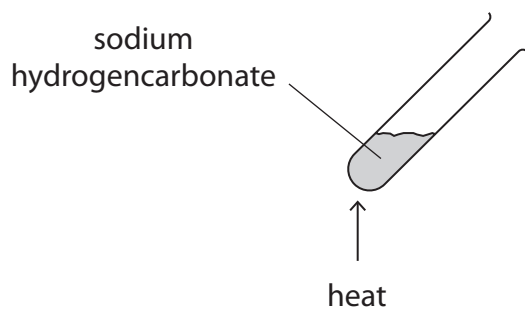
Give your answer to one decimal place. (2)

relative atomic mass =

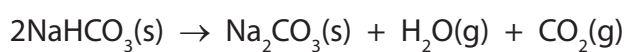
(Total for Question 2 = 5 marks)



3 A student uses this apparatus to investigate the action of heat on sodium hydrogencarbonate (NaHCO_3).



The equation for the reaction is



(a) (i) State the type of reaction taking place.

(1)

(ii) Describe a test to show that the gas given off is carbon dioxide.

(2)

test.....

result.....

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(b) The student heats a 1.00 g sample of sodium hydrogencarbonate for one minute.

He then measures the mass of solid left in the test tube.

He repeats the experiment four times, heating separate samples of mass 1.00 g for a different number of minutes each time.

The table shows the student's results.

Time in minutes	1	2	3	4	5
Mass of solid left in test tube in g	0.89	0.78	0.69	0.63	0.63

(i) State why the mass of solid in each test tube decreases.

(1)

.....

.....

(ii) Suggest why the mass of solid stops decreasing after four minutes.

(1)

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.....

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(Total for Question 3 = 5 marks)



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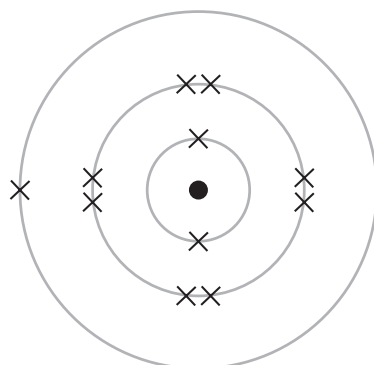
4 Sodium reacts with fluorine to form sodium fluoride.

The reaction is very exothermic.

(a) State what is meant by the term **exothermic**.

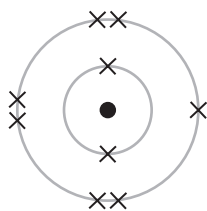
(1)

(b) The diagram shows the electronic configuration of a sodium atom.

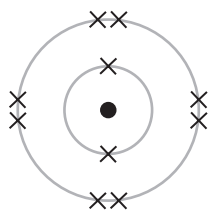


Which of these diagrams shows the electronic configuration of a fluorine atom?

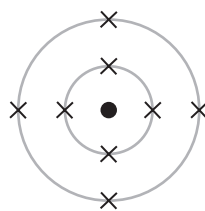
(1)



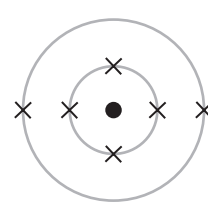
A



B



C



D

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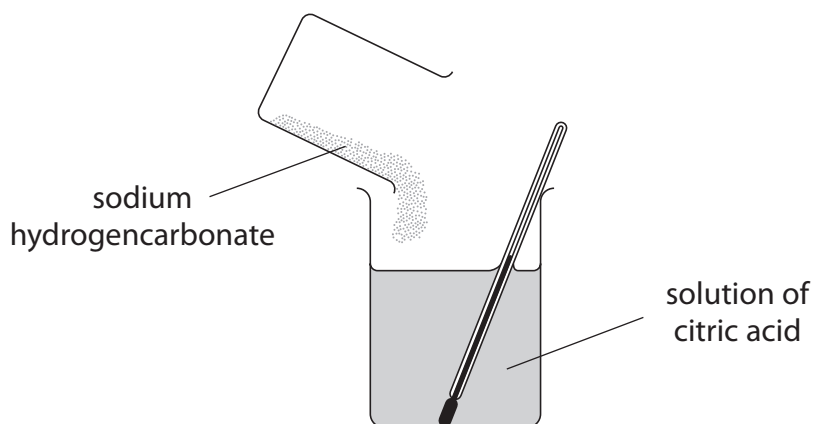
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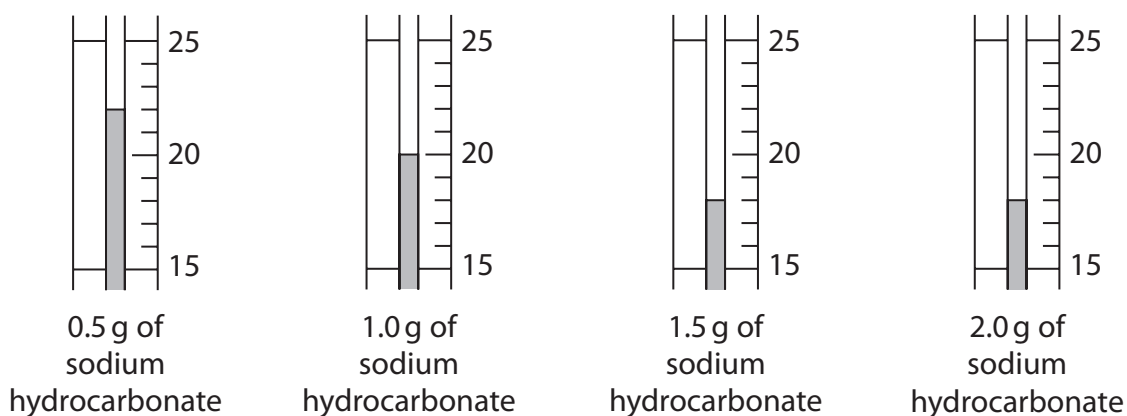


5 A student finds the temperature change when a mass of 0.5 g of sodium hydrogencarbonate is added to 50 cm³ of a solution of citric acid.

She repeats the experiment using masses of 1.0 g, 1.5 g and 2.0 g of sodium hydrogencarbonate.



(a) The diagrams of the thermometer show the lowest temperature reached, in °C, for each experiment.



Use the diagrams to complete the table of results.

(2)

Mass of sodium hydrogencarbonate in g	Initial temperature in °C	Lowest temperature reached in °C	Decrease in temperature in °C
0.5	25		
1.0	24		
1.5	23		
2.0	23		

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(b) Another student does the experiment.

The table shows his results.

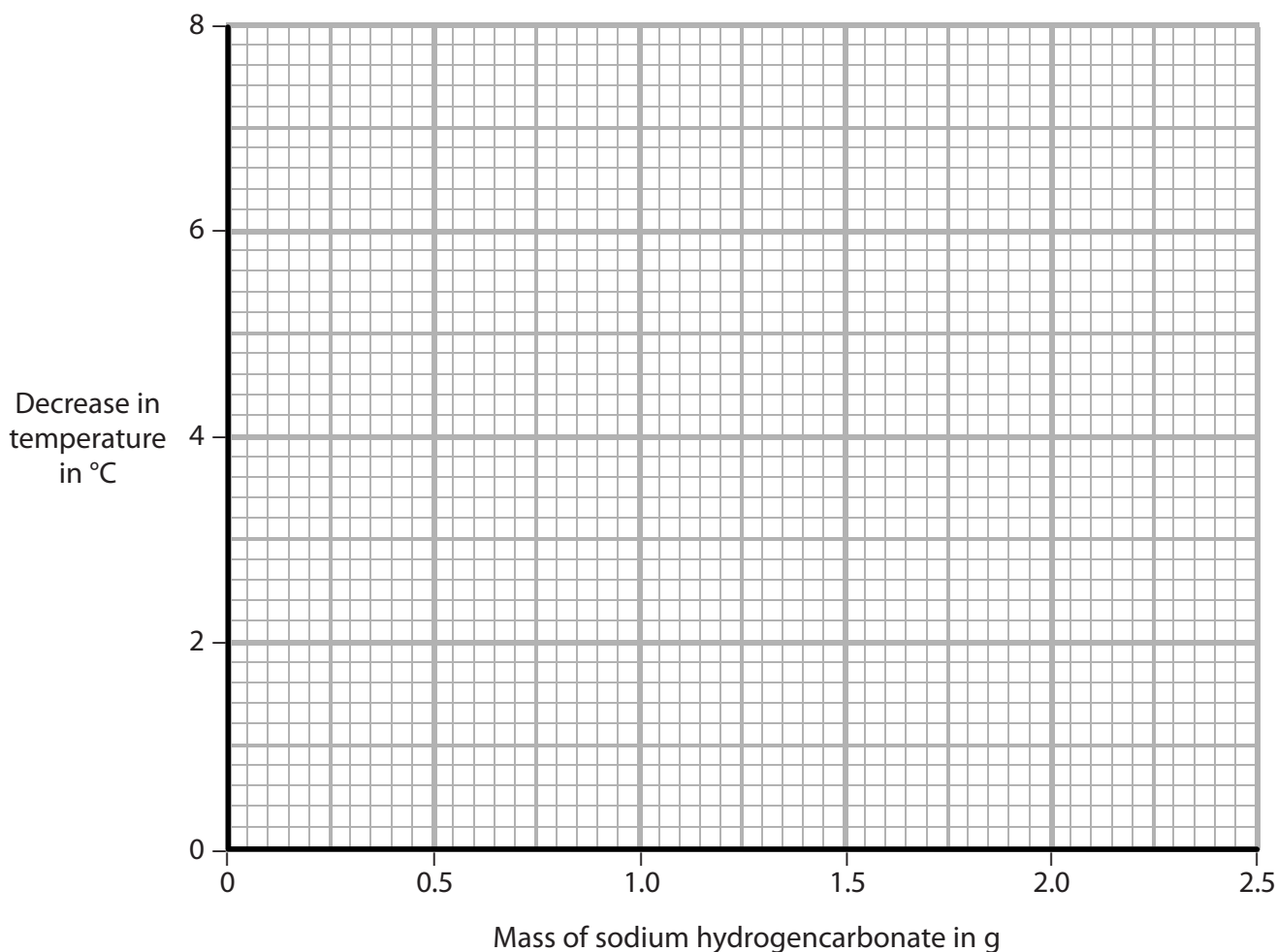
Mass of sodium hydrogencarbonate in g	0.5	1.0	1.5	2.0	2.5
Decrease in temperature in °C	2	4	6	6	6

(i) Plot this student's results on the grid.

Draw a straight line of best fit through the first three points and another straight line of best fit through the last two points.

Make sure the two lines cross.

(3)



(ii) Use your graph to find the mass of sodium hydrogencarbonate required to produce a decrease in temperature of 3°C.

(1)

mass = g

(Total for Question 5 = 6 marks)



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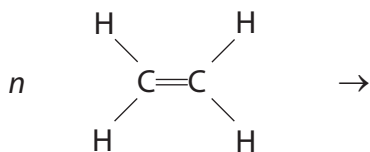
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6 Poly(ethene) is an addition polymer made from ethene, C₂H₄

(a) Complete the equation to show the formation of poly(ethene) from ethene.

(2)



(b) State why poly(ethene) is described as an addition polymer, not a condensation polymer.

(1)

.....

.....

(c) Many shopping bags are made of poly(ethene).

(i) One useful property of poly(ethene) is that it is inert so it does not react with food.

Explain two other properties of poly(ethene) that make it useful for shopping bags.

(2)

1

.....

2

.....

(ii) Another property of poly(ethene) is that it is non-biodegradable.

Two methods of disposing of poly(ethene) are landfill and burning.

Give one problem caused by each method of disposal.

(2)

landfill

.....

burning

.....

(Total for Question 6 = 7 marks)

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7 Magnesium can be obtained by the electrolysis of magnesium chloride.

Solid magnesium chloride is obtained from seawater.

The magnesium chloride is melted and then electrolysed. The positive electrode is made of graphite and the negative electrode is made of steel.

Magnesium forms at the negative electrode. Chlorine forms at the positive electrode.

- (a) Explain why the magnesium chloride has to be melted before it can be electrolysed. (2)

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- (b) Write an ionic half-equation to represent the formation of magnesium at the negative electrode. (1)

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- (c) Suggest why steel is **not** used for the positive electrode. (1)

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(Total for Question 7 = 4 marks)



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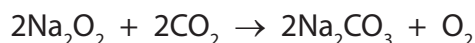
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- 8 Submarines that spend a long time underwater use sodium peroxide (Na_2O_2) to absorb carbon dioxide (CO_2) from the air in the submarine.

The equation for the reaction is



- (a) There are 140 people on the submarine.

Each person produces 480 dm^3 of carbon dioxide per day.

- (i) Calculate the total amount, in moles, of carbon dioxide produced on the submarine in one day.

[assume 1 mol of CO_2 occupies 24.0 dm^3]

(2)

amount of $\text{CO}_2 = \dots\dots\dots$ mol

- (ii) Calculate the mass, in kilograms, of sodium peroxide required to absorb all of the carbon dioxide produced in the submarine in one day.

[M_r of $\text{Na}_2\text{O}_2 = 78.0$]

(2)

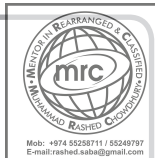
mass of $\text{Na}_2\text{O}_2 = \dots\dots\dots$ kg

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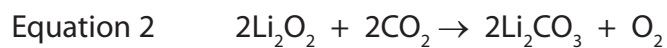
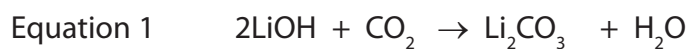
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(b) Spaceships use either lithium hydroxide (LiOH) or lithium peroxide (Li₂O₂) to absorb carbon dioxide.

The equations for the two reactions are



Using information from the equations, give two reasons why lithium peroxide is more suitable than lithium hydroxide for use on spaceships.

(2)

1

.....

2

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(Total for Question 8 = 6 marks)



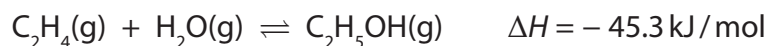
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9 Ethanol (C₂H₅OH) is made in industry by reacting ethene (C₂H₄) with steam at a temperature of 300°C and a pressure of 70 atm. The percentage yield of ethanol is 43%.

The equation for the reaction is



(a) (i) State what the symbols \rightleftharpoons and ΔH represent.

(2)

\rightleftharpoons

ΔH

(ii) Name the catalyst used in this industrial process.

(1)

.....

(b) (i) Predict the effect on the yield of ethanol if the reaction is carried out at a temperature lower than 300°C, but at the same pressure of 70 atm.
[assume reaction reaches equilibrium]

Give a reason for your answer.

(2)

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(ii) Predict the effect on the yield of ethanol if the reaction is carried out at a pressure lower than 70 atm, but at the same temperature of 300°C.
[assume reaction reaches equilibrium]

Give a reason for your answer.

(2)

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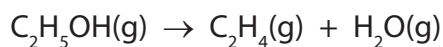
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(c) One method of obtaining ethene is by cracking crude oil fractions.

Ethene can also be made by passing ethanol vapour over a hot aluminium oxide catalyst.

The equation for the reaction is



(i) State the type of reaction taking place.

(1)

(ii) Suggest why it may be necessary, in the future, to make ethene using this reaction rather than by cracking crude oil fractions.

(1)

(Total for Question 9 = 9 marks)



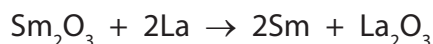
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10 Samarium, Sm, is a metal used to make powerful magnets.

(a) Samarium can be obtained by heating its oxide with lanthanum, La.



The table shows the melting points of the substances involved in this reaction.

Substance	samarium	samarium oxide	lanthanum	lanthanum oxide
Melting point in °C	1072	2335	920	2315

(i) The operating temperature for this reaction is 1030 °C.

Explain which substance in the table could exist as a liquid at this temperature.

(2)

.....

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.....

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(ii) Samarium oxide neutralises hydrochloric acid to form samarium chloride, SmCl_3

Write a chemical equation for this reaction.

(1)

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